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The Ideal Gas Law, continued
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•The ideal gas law describes the behavior of real gases quite well at room temperature and atmospheric pressure.

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One iron atom has a mass of 55.847 amu, and 55.847 g of iron contains 6.022 137 10²³ atoms of iron. Likewise, 32.066 g of sulfur contains 6.022 137 10²³ atoms of sulfur. Knowing this, you can measure out 55.847 g of iron and 32.066 g of sulfur and be pretty certain that you have the same number of atoms of each.

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